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(CHEM.A.2.3) Purpose: In this experiment, you will use your knowledge of periodic properties and a list of clues to correctly arrange the



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elements from a scrambled periodic table. You will also predict values for any information missing from the table. Theory: The present organization of the elements is a product of the first periodic table published by Dmitri Mendeleev in 1869.

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Answers; The periodic law was developed independently by Dmitri Mendeleev and Lothar Meyer in 1869. Mendeleev created the first periodic table and was shortly followed by

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Meyer. They both arranged the elements by their mass and proposed that certain properties periodically reoccur. Meyer formed his periodic law based on the atomic volume or molar volume, which is the atomic mass divided by the density in solid form.

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key features of the periodic table. 2.  
Explain why elements in the same

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group have similar properties. 3. Identify the four blocks of the periodic table based on electron configuration. 4. Compare period and group trends of several properties. 5. Relate period and group trends in atomic radii to electron configuration. Discovery

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Chapter 6: The Periodic Table and Periodic Law

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