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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8)
2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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Chemistry Chapter 8 Covalent Bonding. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. adamjgillman. 8.1 Molecular Compounds 8.2 The Nature of Covalent Bonding 8.3 Bonding Theories 8.4 Polar Bonds and Molecules. Terms in this set (31) Bond dissociation energy.

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## 8: Covalent Bonding and Molecular Structure

covalent bonds. The majority of covalent bonds form between nonmetallic elements. 8. Describe how the octet rule applies to covalent bonds. Atoms share valence electrons; the shared electrons complete the octet of each atom. 9. Illustrate the



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formation of single, double, and triple covalent bonds using Lewis structures.

Covalent Bonding  
Covalent Bonding

Covalent Bonding □ There are several electrostatic interactions in these bonds: - Attractions between electrons and nuclei - Repulsions between electrons - Repulsions between nuclei □

Covalent bond, sharing electrons, □ But electron sharing not always equal.

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This video describes how atoms  
covalently bond, and form single,  
double or triple bonds. Pi bonds  
are discussed as well as bond  
strength.

Chapter 8 Covalent Bonding Pt 1  
Valence bond theory describes a  
covalent bond as the overlap of

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half-filled atomic orbitals (each containing a single electron) that yield a pair of electrons shared between the two bonded atoms. We say that orbitals on two different atoms overlap when a portion of one orbital and a portion of a second orbital occupy the same region of space.

## 8.1 Valence Bond Theory | Chemistry

A covalent bond is formed between atoms held together by sharing electrons. A molecule is a group of atoms joined by covalent bonds. Section 8.2 Assessment. PDF Chapter 8: Covalent Bonding. 242 Chapter 8 □ Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it

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